

Dalton's Atomic Theory

- Elements are composed of atoms.
- Atoms make up the materials we have.
- Chemical reactions make chemicals rearrange structure.
- Atoms mix together in **whole number** ratios to form compounds.

#15: Table

Element	Atomic #	Protons	Electrons
K	19		19
			5
S	16		
V		23	

Definitions

- Atomic #: Number of Protons
- #of Protons = # of Electrons
- **Whole #** on periodic table.

Definitions

- Atomic Mass:
- Mass #:
- AMU:

Number of Neutrons

- Mass # - Atomic # = # of Neutrons.
- Tell me the number of Protons, Neutrons, and Electrons.
- Br-80
- Rb-86

Calculating Atomic Mass

- Weighted or Unweighted Average?
Explain?
- Formula:
- Cw- “Buckium” has 3 isotopes.
Cw-45 (74%), Cw-47 (18%), & Cw-48 (8%)
Determine the Atomic Mass.

- **Cw-45 (74%), Cw-47 (18%), & Cw-48 (8%)**

List the number of protons, neutrons, and electrons for each isotope.

Element	Protons	Neutrons	Electrons
P-31			
P-32			
Zr-92			
Zr-94			